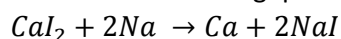


- c. If 30 mol of H₂O is produced, determine the moles of CO₂ produced
- d. If 0.015 mol of CO₂ is produced, determine the moles of C₅H₁₂ reacting

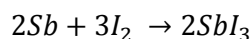
2. Use the following balanced equation to answer the following questions.



- a. If 2.2 mol of calcium iodide reacts, determine the moles of sodium reacting and the moles of each product.
- b. If 0.016 mol of sodium reacts, determine the moles of calcium iodide reacting and the moles of each product.

V. Mass-Mass stoichiometry

1. Antimony and iodine can be combined to form antimony triiodide. If 60.09 g of antimony are present, what mass of iodine will be required in the reaction? What mass of antimony triiodide will be produced?



2. Strontium chloride reacts with sodium oxalate to produce strontium oxalate and sodium chloride. If 3.20 g of strontium chloride are present, determine the mass of sodium oxalate required in the reaction. What mass of strontium oxalate and sodium chloride will be produced?

