Name: _____ Per____

Mole Conversions: Part 2

Practice Sheet #18

1. What is the mass of 8.48 moles of Boron?

2. What is the mass of 8.48 moles of Oxygen?

3. How many moles are in 65.8 g of ammonium sulfate?

4. What is the mass of 3.45×10^{22} molecules of hydrochloric acid?

5. What is the mass of 3.5×10^{15} molecules of bromic acid?

6. How many molecules are in 9.2 grams of BH_3?

7. How many molecules are in 0.0078 grams of sulfur hexafluoride?

8. Find the number of atoms of EACH element in 0.0078 grams of hexafluoride.

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9. What are hydrates?

10. How do you name hydrates?

11. How do your write the formula for hydrates?

12. Write the name for each of the following hydrates.
    a. $Cs_2CO_2 \cdot 2H_2O$
    b. $MgSO_4 \cdot 7H_2O$
    c. $FeCl_3 \cdot 6H_2O$
    d. $Ca(NO_3)_2 \cdot 4H_2O$

13. Write the formula for each of the following hydrates.
    a. Calcium chloride dehydrate
    b. Cobalt (II) chloride hexahydrate
    c. Magnesium hydrogen phosphate trihydrate
    d. Aluminum hypochlorite octahydrate