

Name: _____ Per _____

Mole Conversions: Part 2

Practice Sheet #18

1. What is the mass of 8.48 moles of Boron?
2. What is the mass of 8.48 moles of Oxygen?
3. How many moles are in 65.8 g of ammonium sulfate?
4. What is the mass of 3.45×10^{22} molecules of hydrochloric acid?
5. What is the mass of 3.5×10^{15} molecules of bromic acid?
6. How many molecules are in 9.2 grams of BH_3 ?
7. How many molecules are in 0.0078 grams of sulfur hexafluoride?
8. Find the number of atoms of EACH element in 0.0078 grams of hexafluoride.

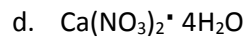
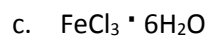
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9. What are hydrates?

10. How do you name hydrates?

11. How do you write the formula for hydrates?

12. Write the name for each of the following hydrates.



13. Write the formula for each of the following hydrates.

a. Calcium chloride dehydrate

b. Cobalt (II) chloride hexahydrate

c. Magnesium hydrogen phosphate trihydrate

d. Aluminum hypochlorite octahydrate