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Practice Multiple Choice Semester 1 Final Exam
Self-Assessment

Directions: Listed below is the number for each multiple choice problem on the practice multiple choice semester 1 final. Next to each number that you got wrong 1. check the box indicated that you got this wrong 2. Rework the problem to get the correct answer. After answer the additional problems

MC #	Unit	Topic	Wrong?	For any questions that you got wrong rework the problem for the correct answer.
1	1	Convert between scientific notation and standard form of a number.		
2	1	Convert between scientific notation and standard form of a number.		
3	1	Convert from one metric unit to another.		
4	1	Perform calculations involving density, mass, and volume giving answer with appropriate units.		
5	2	Use the periodic table to determine the number of protons, electrons, neutrons, and atomic mass for a given element.		
6	2	Use the periodic table to determine the number of protons, electrons, neutrons, and atomic mass for a given element.		
7	2	Determine the number of electrons for an ion.		
8	2	Determine the number of valence electrons for an atom.		
9	2	Give orbital notation for a given atom/ion		

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10	2	Write Electronic configuration for a given atom/ion		
11	2	Write noble gas configuration for a given atom/ion		
12	2	Write Electronic configuration for a given atom/ion		
13	3	Define atomic radius and explain periodic trends in this property as they relate to atomic structure		
14	3	Define ionization energy and explain PT trends in this property as they relate to atomic structure		
15	3	Draw Lewis Structures for molecules from chemical formulas.		
16	3	Draw Lewis Structures for molecules from chemical formulas.		
17	3	Assign shapes to molecules using VSEPR Theory		
18	3	Assign shapes to molecules using VSEPR Theory		
19	4	Name and write formulas for ionic compounds including compounds with multivalent ions and those with polyatomic ions.		
20	4	Name and write formulas for ionic compounds including compounds with multivalent ions and those with polyatomic ions.		

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21	4	Name and write formulas for ionic compounds including compounds with multivalent ions and those with polyatomic ions.		
22	4	Name and write formulas for ionic compounds including compounds with multivalent ions and those with polyatomic ions.		
23	4	Name and write formulas for covalent compounds using prefixes		
24	4	Name and Write formulas for acids.		
25	5	Determine the molar mass for a given element compound with the appropriate units.		
26	5	Perform conversions between moles, mass and atoms/molecules and solve problems involving these quantities.		
27	5	Perform conversions between moles, mass and atoms/molecules and solve problems involving these quantities.		
28	5	Perform conversions between moles, mass and atoms/molecules and solve problems involving these quantities.		
29	5	Perform conversions between moles, mass and atoms/molecules and solve problems involving these quantities.		
30	5	Perform conversions between moles, mass and atoms/molecules and solve problems involving these quantities.		
31	5	Determine the percent composition for a compound from the chemical formula		

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32	6	Classify reactions as synthesis, decomposition, SR, DR, or combustion.		
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34	6	Classify reactions as synthesis, decomposition, SR, DR, or combustion.		
35	6	Classify reactions as synthesis, decomposition, SR, DR, or combustion.		
36	6	Balance chemical equations beginning with either chemical names or formulas.		
37	6	Determine mole ratios for a reaction from the balanced chemical equation in order to convert between moles of different substances.		
38	6	Determine mole ratios for a reaction from the balanced chemical equation in order to convert between moles of different substances.		
39	6	Perform stoichiometry calculations involving mass or a reactant or product.		
40	6	Perform stoichiometry calculations involving mass or a reactant or product.		

Count up the number of questions that you missed for each unit and then fill out the following table.

Unit	# of questions missed
1	
2	
3	
4	
5	
6	

Which unit do you need to study the most? _____