

Name: _____ Per _____

VSEPR Model Building

1. What does VSEPR stand for?
2. What general, what principles determine the shape of a molecule?
3. Which takes up more space, a lone pair of electrons or a bond? Explain.

4. Using VSEPR theory fill out the following chart.

Molecule	Lewis Structure	# atoms bonded to the central atom	# lone pairs	VSEPR Diagram	Shape	Polar or Non-Polar
CCl₄						
BH₃						
NI₃						
SH₆						
SeCl₂						
PI₅						
SiS₂						

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* Bonus Question(1 point extra credit): In the above glycine structure determine the shape of each central atom.

Central Atom	# atoms bonded to Central atom	# lone pairs on central atom	VSEPR Shape
Nitrogen			
Carbon 1			
Carbon 2			
Oxygen			

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