

Percent Yield

Practice Sheet #

1. Strontium nitride can be decomposed into strontium and nitrogen.
 - a. Write a balanced equation for this reaction.

 - b. If 58.2 g of strontium nitride yields 45.0 g of strontium, determine the percent yield for the reaction.

 - c. What mass of nitrogen will actually be obtained?

2. Mercury (II) oxide can be decomposed into mercury and oxygen.
 - a. Write a balanced chemical equation for this reaction.

 - b. If 54.15 g of mercury (II) oxide yields 3.250 g of oxygen, determine the percent yield for the reaction.

 - c. What mass of mercury will actually be obtained?

3. Zinc and fluorine can be combined to form zinc fluoride.
 - a. Write a balanced chemical equation for this reaction.

Name: _____ Per _____

- b. If 16.35 g of zinc yields 16.50 g of zinc fluoride, determine the percent yield of this reaction.
 - c. If 0.380 g of fluorine reacts with a percent yield of 60.0 %, what mass of zinc fluoride will actually be obtained?

4. Magnesium chloride reacts with silver nitrate.
 - a. Write a balanced chemical equation for this reaction.
 - b. If 21.2 g of silver nitrate react with a percent yield of 80.0 %, what mass would actually be obtained for each of the products?
 - c. If 476 g of magnesium chloride yields 400 g of magnesium nitrate, determine the percent yield of the reaction. What mass of silver chloride would actually be obtained?

5. Aluminum is reacted with copper (II) sulfate.
 - a. Write a balanced chemical equation for this reaction.
 - b. If 5.60 g of aluminum react, determine the mass of copper (II) sulfate required.
 - c. If 18.0 g of copper are actually produced, determine the percent yield of the reaction. What mass of aluminum sulfate will actually be obtained?