

Name: \_\_\_\_\_ Per\_

**Dissociation**  
*Practice Sheet #34*

1. Write the dissociate equations for the following compounds.
  - a. HI
  - b.  $\text{CaBr}_2$
  - c.  $\text{MgSO}_4$
  - d.  $\text{Al}(\text{NO}_3)_3$
  - e.  $\text{CuCl}_2$
  - f. KCN
  - g.  $\text{Cr}_2(\text{SO}_4)_3$
  - h.  $(\text{NH}_4)_2\text{SO}_4$
  - i. Sodium carbonate
  - j. Cobalt (II) chloride
  
2. Write the dissociation equation for each of the following compounds. Calculate the concentration of each ion in solution.
  - a. 2.0 M NaF
  
  - b. 0.50 M  $\text{BaI}_2$
  
  - c. 0.40 M  $\text{K}_3\text{PO}_4$
  
  - d. 0.080 M  $\text{CrCl}_3$
  
  - e. 0.10 M  $\text{Ni}(\text{NO}_3)_2$
  
  - f.  $2.0 \times 10^{-3}$  M  $\text{ZnBr}_2$
  
  - g. 0.050 M  $\text{K}_2\text{Cr}_2\text{O}_7$

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3. Calculate the concentration of each ion resulting from mixing the following solutions, given that no reaction occurs.

a. 4.0 L of 1.2 M HBr mixed with 8.0 L of 2.4 M NaCl

b. 2.0 L of 0.25 M  $\text{KNO}_3$  mixed with 3.0 L of 0.45 M  $\text{SrI}_2$

c. 150 mL of 0.80 M  $\text{CaCl}_2$  mixed with 50 mL of 0.60 M  $\text{Mg}(\text{NO}_3)_2$

d. 0.40 L of 0.35 M NaBr mixed with 0.60 L of 0.20 M  $\text{Na}_2\text{CO}_3$

e. 20 mL of 0.15 M  $\text{Al}(\text{NO}_3)_3$  mixed with 10 mL of 0.30 M  $\text{Pb}(\text{NO}_3)_2$