

Name: _____ per _____

Solubility
Practice Sheet # 35

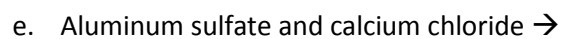
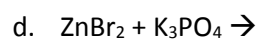
1. Determine whether each of the following compounds is soluble or insoluble in water.
 - a. CaF₂
 - b. Sodium hydroxide
 - c. Iron (II) carbonate
 - d. PbBr₂
 - e. Ca(OH)₂
 - f. Silver chloride
 - g. Ammonium phosphate
 - h. K₂S
 - i. CuI₂
 - j. Calcium sulfate

2. Write the formula equation, complete ionic equation, and the net ionic equation for the following reactions.
 - a. Silver nitrate and sodium chloride

 - b. $\text{MgSO}_4 + \text{KF} \rightarrow$

 - c. $\text{FeCl}_3 + \text{NaOH} \rightarrow$

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3. Describe a procedure to separate each of these ions from solution.

a. Cl^- and F^-

b. Ag^+ and Cu^{2+}

c. Cu^+ , Mg^{2+} , and Sr^{3+}

d. SO_4^{2-} , OH^- , and S^{2-}

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e. Ba^{2+} , Ag^+ , and Ca^{2+}